



# Cambridge IGCSE™

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## CHEMISTRY

0620/33

Paper 3 Theory (Core)

October/November 2023

1 hour 15 minutes

You must answer on the question paper.

No additional materials are needed.

### INSTRUCTIONS

- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working and use appropriate units.

### INFORMATION

- The total mark for this paper is 80.
- The number of marks for each question or part question is shown in brackets [ ].
- The Periodic Table is printed in the question paper.

This document has **20** pages. Any blank pages are indicated.

1 A list of substances is shown.

ammonia  
calcium oxide  
carbon monoxide  
cobalt(II) chloride  
ethane  
ethanol  
ethene  
oxygen  
potassium oxide  
sodium sulfate  
sulfuric acid  
water

Answer the following questions using only the substances from the list.  
Each substance may be used once, more than once or not at all.

Give the name of the substance that:

(a) is a product of photosynthesis

..... [1]

(b) is a member of the alkene homologous series

..... [1]

(c) has an ion with a charge of 1–

..... [1]

(d) is used to remove sulfur dioxide in flue gas desulfurisation

..... [1]

(e) is the product formed in a hydrogen–oxygen fuel cell

..... [1]

(f) is used to test for water.

..... [1]

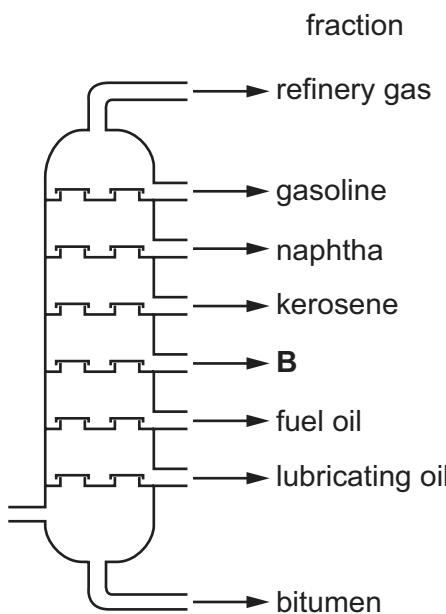
[Total: 6]

2 Hydrocarbons are compounds of carbon and hydrogen.

(a) State the meaning of the term compound.

..... [2]

(b) Fig. 2.1 shows a fractionating column for separating petroleum into different hydrocarbon fractions.



**Fig. 2.1**

(i) On Fig. 2.1, draw an **X** inside the column to show where the hydrocarbon with the highest boiling point collects. [1]

(ii) Name the fraction labelled **B** in Fig. 2.1.

..... [1]

(iii) State the name of the fraction which has hydrocarbons with the shortest chain length.

..... [1]

(iv) State **one** use of the naphtha fraction.

..... [1]

[Total: 6]

3 (a) Table 3.1 shows the average concentrations, in ng/1000 cm<sup>3</sup>, of air pollutants in four different years.

**Table 3.1**

year	concentration of air pollutant in ng/1000 cm <sup>3</sup>				
	carbon monoxide	hydrocarbons	oxides of nitrogen	particulates	sulfur dioxide
2019	5.3	22.0	15.6	19.0	20.0
2020	4.1	13.5	14.8	20.1	18.2
2021	5.8	14.8	22.7	23.5	16.2
2022	2.6	18.0	10.9	26.2	14.0

(i) Name the pollutant which has the highest concentration in 2019.

..... [1]

(ii) Name the pollutant that shows a continuous decrease in concentration from 2019 to 2022.

..... [1]

(iii) Calculate the average mass, in ng, of hydrocarbons in a 200 cm<sup>3</sup> sample of polluted air in 2019.

mass = ..... ng [1]

(b) (i) State **one** source of oxides of nitrogen in the air.

..... [1]

(ii) Oxides of nitrogen contribute to acid rain.

Give one **other** effect of oxides of nitrogen in the air.

..... [1]

(iii) Unpolluted water has a neutral pH.

Choose from the list the pH value of a neutral substance.

Draw a circle around your chosen answer.

pH1

pH6

pH7

pH14

[1]

(c) Nitrogen dioxide is an acidic oxide.

Choose an oxide from the list which is also an acidic oxide.

Tick ( $\checkmark$ ) **one** box.

copper(II) oxide	<input type="checkbox"/>
magnesium oxide	<input type="checkbox"/>
phosphorus(V) oxide	<input type="checkbox"/>
sodium oxide	<input type="checkbox"/>

[1]

(d) Sulfur dioxide reacts with oxygen to produce sulfur trioxide.

(i) Complete the symbol equation for this reaction.



[2]

(ii) State the meaning of the symbol  $\rightleftharpoons$ .

..... [1]

(iii) Sulfur trioxide reacts with calcium oxide to produce calcium sulfate.

Describe a test for sulfate ions.

test .....

observations .....

[2]

[Total: 12]

4 Nitrogen is a gas at room temperature.

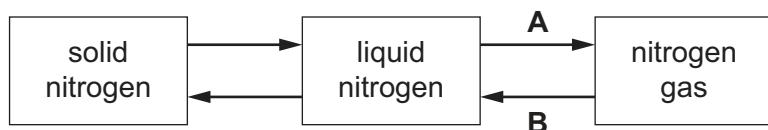
(a) State **two** general properties of a gas.

1 .....

2 .....

[2]

(b) Fig. 4.1 shows the physical states of nitrogen.



**Fig. 4.1**

Name the changes of physical states **A** and **B**.

**A** .....

**B** .....

[2]

(c) Describe solid nitrogen and nitrogen gas in terms of the arrangement and separation of the particles.

solid nitrogen

arrangement .....

separation .....

nitrogen gas

arrangement .....

separation .....

[4]

(d) A sealed gas syringe contains  $80\text{ cm}^3$  of nitrogen gas.

State how increasing the pressure affects the volume of nitrogen gas in the gas syringe when the temperature remains constant.

..... [1]

[Total: 9]

5 This question is about metals.

(a) Table 5.1 shows some properties of the Group I metals.

**Table 5.1**

metal	melting point in °C	boiling point in °C	atomic volume in cm <sup>3</sup> /mol	observations on reaction with water
lithium	181	1342	12.9	bubbles form slowly but no flame
sodium	98	883	23.7	
potassium	63	760		bubbles form very rapidly and flame seen
rubidium		686	55.8	explodes

Use the information in Table 5.1 to predict:

(i) the melting point of rubidium ..... [1]

(ii) the atomic volume of potassium ..... [1]

(iii) the observations when sodium reacts with water .....  
..... [1]

(iv) the physical state of sodium at 1300 °C. Give a reason for your answer.

physical state .....

reason .....

..... [2]

(b) Iron is extracted in a blast furnace by reduction of iron(III) oxide.

(i) In the first step, carbon burns in air to form carbon dioxide.

State the percentage of oxygen in clean, dry air.

..... [1]

(ii) In the second step, carbon monoxide is produced by the reaction of carbon dioxide with carbon.



Choose the correct statement about this reaction.

Tick (✓) one box.

the carbon dioxide is oxidised and the carbon is reduced

both carbon dioxide and carbon are oxidised

the carbon dioxide is reduced and the carbon is oxidised

both carbon dioxide and carbon are reduced

[1]

(iii) In the third step, iron(III) oxide is reduced by carbon monoxide.

The reaction is exothermic.

State the meaning of the term exothermic.

..... [2]

(c) Calcium carbonate is added to the blast furnace.

The calcium carbonate breaks down as shown.



(i) Name the type of chemical reaction that takes place.

..... [1]

(ii) Complete this sentence about the calcium oxide that is produced in the blast furnace.

Calcium oxide reacts with impurities in the iron ore to form ..... [1]

(d) Table 5.2 gives the observations when four different metals react with air.

**Table 5.2**

metal	observations
cerium	forms an oxide layer slowly without heating
copper	forms an oxide layer only when heated
gold	does not form an oxide layer even when heated
rubidium	forms an oxide layer quickly without heating

Put the four metals in order of their reactivity.

Put the least reactive metal first.

least reactive  most reactive

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[2]

[Total: 13]

6 Hydrogen peroxide,  $\text{H}_2\text{O}_2$ , breaks down slowly at 40 °C to produce oxygen gas and water.



A student investigates the breakdown of hydrogen peroxide at 40 °C in the presence of a catalyst.

(a) Fig. 6.1 shows the volume of oxygen gas released as the reaction proceeds.

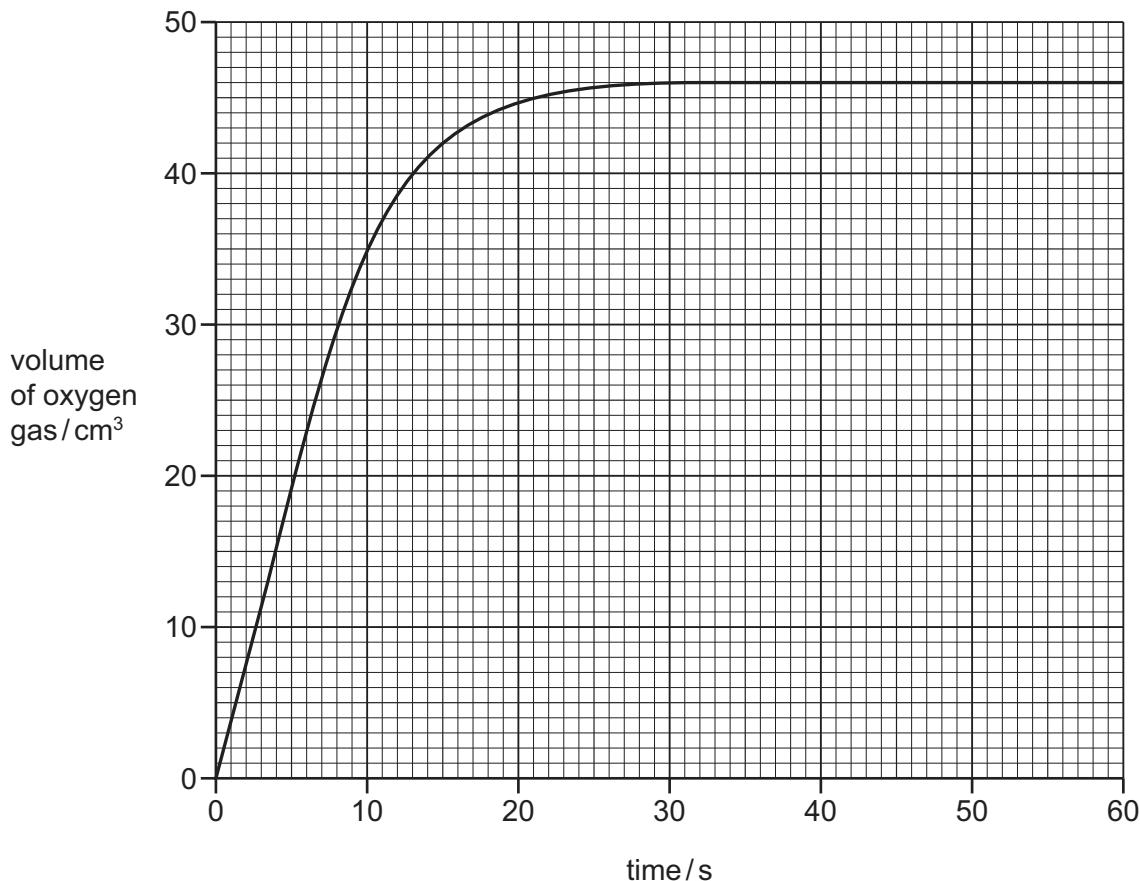


Fig. 6.1

(i) Deduce the volume of oxygen gas released after 15 seconds.

$$\text{volume of oxygen} = \dots \text{cm}^3 \quad [1]$$

(ii) The student repeats the experiment at 20 °C.

All other conditions stay the same.

Draw a line on the grid in Fig. 6.1 to show how the volume of oxygen changes when a temperature of 20 °C is used. [2]

(b) (i) The student repeats the experiment without a catalyst.

All other conditions stay the same.

Describe how the rate of reaction differs when no catalyst is used.

..... [1]

(ii) The student repeats the experiment using a lower concentration of hydrogen peroxide.

All other conditions stay the same.

Describe how the rate of reaction differs when a lower concentration of hydrogen peroxide is used.

..... [1]

(c) Hydrogen peroxide can act as a reducing agent in the presence of an alkali.

(i) State the meaning of the term alkali.

..... [1]

(ii) Give the formula of the ion that is present in all alkaline solutions.

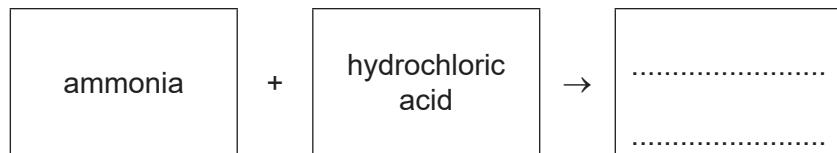
..... [1]

(iii) State the colour of methyl orange in an alkaline solution.

..... [1]

(iv) Aqueous ammonia is an alkali.

Complete the word equation for the reaction of aqueous ammonia with hydrochloric acid.

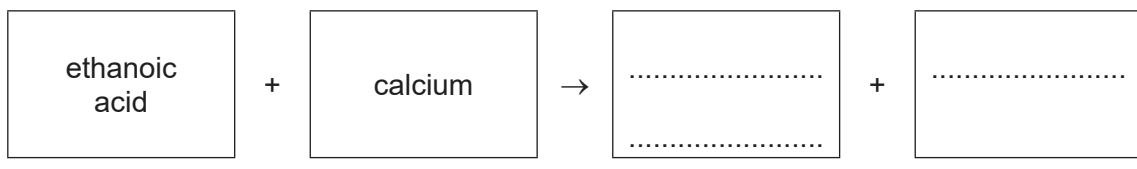


[1]

[Total: 9]

7 Ethanoic acid and methacrylic acid are both carboxylic acids.

(a) Complete the word equation for the reaction of ethanoic acid with calcium.



[2]

(b) Ethanoic acid can be reduced to ethanol.

(i) Name the homologous series that includes ethanol.

..... [1]

(ii) Ethanol can be manufactured by fermentation.

Describe **two** conditions needed for fermentation.

1 .....

2 .....

[2]

(c) Fig. 7.1 shows the displayed formula of methacrylic acid.

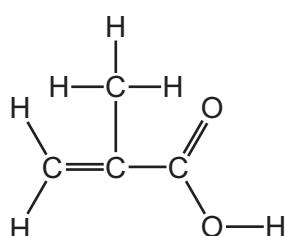


Fig. 7.1

(i) On Fig. 7.1, draw a circle around the functional group which reacts with aqueous bromine. [1]

(ii) State the colour of aqueous bromine.

..... [1]

(iii) Deduce the molecular formula of methacrylic acid.

..... [1]

(d) Methacrylic acid can be converted to methyl methacrylate.  
The molecular formula of methyl methacrylate is  $C_5H_8O_2$ .

Complete Table 7.1 to calculate the relative molecular mass of methyl methacrylate.

**Table 7.1**

atom	number of atoms	relative atomic mass	
carbon	5	12	$5 \times 12 = 60$
hydrogen		1	
oxygen		16	

relative molecular mass = ..... [2]

(e) Methyl methacrylate can be polymerised to produce a plastic.

Describe **two** environmental problems caused by plastics.

1 .....

2 .....

[2]

(f) Poly(ethene) is a polymer.

Draw the displayed formula of the monomer used to make poly(ethene).

[1]

[Total: 13]

8 Potassium chloride is an ionic compound.

(a) Complete Fig. 8.1 to show:

- the electronic configuration of a potassium ion
- the charge on the ion.

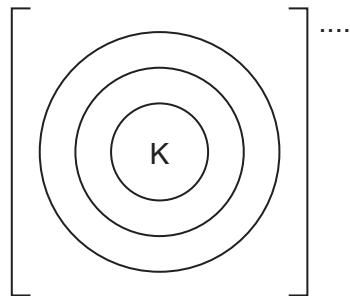


Fig. 8.1

[2]

(b) Deduce the number of protons and neutrons in the chloride ion shown.



number of protons .....

number of neutrons .....

[2]

(c) Molten potassium chloride is electrolysed using graphite electrodes.

(i) Define the term electrolysis.

.....  
..... [2]

(ii) State the names of the products at each electrode and give the observations at the positive electrode.

product at the negative electrode .....

product at the positive electrode .....

observations at the positive electrode

.....  
..... [3]

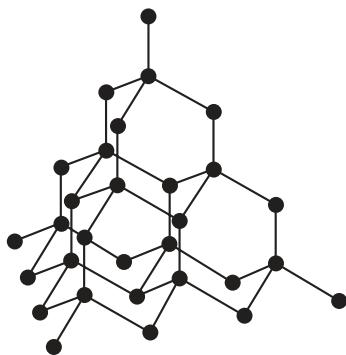
(d) Graphite electrodes are inert.

Name one **other** inert electrode.

..... [1]

(e) Graphite and diamond are two forms of carbon.

Fig. 8.2 shows the structure of diamond.



**Fig. 8.2**

(i) Name the type of bonding in diamond.

..... [1]

(ii) Use Fig. 8.2 to explain why diamond is used in cutting tools.

..... [1]

[Total: 12]





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# The Periodic Table of Elements

Group		Group																				
		I		II		III		IV		V		VI		VII								
3	4	<b>Li</b> lithium 7	<b>Be</b> beryllium 9	<b>Sc</b> scandium 45	<b>Ti</b> titanium 48	<b>V</b> vanadium 51	<b>Cr</b> chromium 52	<b>Mn</b> manganese 55	<b>Fe</b> iron 56	<b>Co</b> cobalt 59	<b>Ni</b> nickel 59	<b>Cu</b> copper 64	<b>Zn</b> zinc 65	<b>Ga</b> gallium 70	<b>Ge</b> germanium 73	<b>As</b> arsenic 75	<b>Se</b> selenium 79	<b>Br</b> bromine 80	<b>Kr</b> krypton 84			
11	12	<b>Na</b> sodium 23	<b>Mg</b> magnesium 24	<b>Ca</b> calcium 40	<b>Y</b> yttrium 89	<b>Zr</b> zirconium 91	<b>Nb</b> niobium 93	<b>Mo</b> molybdenum 96	<b>Tc</b> technetium —	<b>Ru</b> ruthenium 101	<b>Rh</b> rhodium 103	<b>Pd</b> palladium 106	<b>Ag</b> silver 108	<b>Cd</b> cadmium 112	<b>In</b> indium 115	<b>Sn</b> tin 119	<b>Te</b> tellurium 122	<b>I</b> iodine 128	<b>Xe</b> xenon 131			
19	20	<b>K</b> potassium 39	<b>Sr</b> strontium 88	<b>Rb</b> rubidium 85	<b>Yt</b> yttrium 89	<b>Zr</b> zirconium 91	<b>Nb</b> niobium 93	<b>Mo</b> molybdenum 96	<b>Tc</b> technetium —	<b>Ru</b> ruthenium 101	<b>Rh</b> rhodium 103	<b>Pt</b> platinum 106	<b>Au</b> gold 197	<b>Tl</b> thallium 195	<b>Pb</b> lead 207	<b>Bi</b> bismuth 209	<b>Po</b> polonium —	<b>At</b> astatine —	<b>Rn</b> radon —			
37	38	<b>Cs</b> caesium 133	<b>Ba</b> barium 137	<b>Fr</b> francium —	<b>La</b> lanthanum 139	<b>Th</b> thorium 232	<b>Pa</b> protactinium 231	<b>U</b> uranium 238	<b>Ta</b> tantalum 181	<b>W</b> tungsten 184	<b>Re</b> rhodium 186	<b>Os</b> osmium 190	<b>Ir</b> iridium 192	<b>Pt</b> platinum 195	<b>Au</b> gold 197	<b>Hg</b> mercury 201	<b>Tl</b> thallium 204	<b>Pb</b> lead 207	<b>Bi</b> bismuth 209	<b>Po</b> polonium —	<b>At</b> astatine —	<b>Rn</b> radon —
55	56	<b>Ca</b> calcium 133	<b>Rf</b> actinoids 89–103	<b>Db</b> rutherfordium —	<b>Pr</b> praseodymium 140	<b>Pa</b> protactinium 231	<b>Pm</b> promethium —	<b>Am</b> americium —	<b>Sm</b> samarium 150	<b>Eu</b> europium 152	<b>Gd</b> gadolinium 157	<b>Tb</b> terbium 159	<b>Dy</b> dysprosium 163	<b>Ho</b> holmium 165	<b>Er</b> erbium 167	<b>Tm</b> thulium 169	<b>Yb</b> ytterbium 173	<b>Lu</b> lutetium 175				
87	88	<b>Ra</b> radium —	<b>Ac</b> actinium —	<b>Fr</b> francium —	<b>Pr</b> cerium 140	<b>Th</b> thorium 232	<b>Pa</b> protactinium 231	<b>Pu</b> plutonium —	<b>Cm</b> curium —	<b>Bk</b> berkelium —	<b>Cf</b> californium —	<b>Fm</b> fermium —	<b>Md</b> mendelevium —	<b>No</b> nobelium —	<b>Os</b> osmium —	<b>Ts</b> tennessine —	<b>Og</b> oganesson —	<b>Fr</b> francium —	<b>Ac</b> actinium —			

<b>lanthanoids</b>	57	58	<b>Ce</b> cerium 140	59	60	<b>Pr</b> praseodymium 141	61	<b>Nd</b> neodymium 144	62	<b>Sm</b> samarium 150	63	<b>Eu</b> europium 152	64	<b>Gd</b> gadolinium 157	65	<b>Tb</b> terbium 159	66	<b>Dy</b> dysprosium 163	67	<b>Ho</b> holmium 165	68	<b>Er</b> erbium 167	69	<b>Tm</b> thulium 169	70	<b>Yb</b> ytterbium 173	71	<b>Lu</b> lutetium 175
<b>actinoids</b>	89	90	<b>Th</b> thorium 232	91	92	<b>Pa</b> protactinium 231	93	<b>Np</b> neptunium —	94	<b>Am</b> americium —	95	<b>Pu</b> plutonium —	96	<b>Cm</b> curium —	97	<b>Bk</b> berkelium —	98	<b>Cf</b> californium —	99	<b>Fm</b> fermium —	100	<b>Md</b> mendelevium —	101	<b>No</b> nobelium —	102	<b>Os</b> osmium —	103	<b>Fr</b> francium —

The volume of one mole of any gas is  $24 \text{ dm}^3$  at room temperature and pressure (r.t.p.).